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Problems. Answer
for Problem # 1.
Since the
containers are
insulated, no
heat transfer
occurs between
the gas and the
external

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environment, and since the gas expands freely into container B there is no resistance "pushing" against it, which means no work is done on the gas as it expands.

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Problem : Given that the free energy of formation of liquid water is -237 kJ / mol , calculate the potential for the formation of hydrogen and oxygen from

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water. To solve this problem we must first calculate ΔG for the reaction, which is $-2 (-237 \text{ kJ / mol}) = 474 \text{ kJ / mol}$. Knowing that $\Delta G = -nFE^\circ$ and $n = 4$, we calculate the potential is -1.23 V .

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Defining key

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Answer The van't Hoff equation is so the slope of the $\ln K$ eq versus $1/T$ plot equals $-\Delta H^\circ/R$.

The slope of the plot = 22700 so $\Delta H^\circ = -R \times \text{slope} =$

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$$\begin{aligned} &= -(8.314) \times (22700) \\ &= -189000 \text{ J/mol} \\ &= -189 \text{ kJ/mol} \end{aligned}$$

CHM 112

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Heat capacity C of a body as the ratio of the amount of heat energy Q transferred to a body in any process to its corresponding temperature change ΔT . $C = Q/\Delta T$. So, $Q = C \Delta T$. Each species will experience

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the equal
temperature
change. If the
gas has n
molecules, then
 Q will be, $Q =$
 $nC_v \Delta T$.

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are common thermodynamic equations and sample problems showing a situation in which each might be used.

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Part 1 Problem

1. A rapidly spinning paddle wheel raises the temperature of 200mL of water from 21 degrees Celsius to 25 degrees. How much a) work is done and b) heat is transferred in this process?
Solution .

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Problem 2. The temperature of a body is increased from -173 C to 357 C .

Physics

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Define the First Law of

Thermodynamics.

Thermal energy can change form

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and location, but it cannot be created or destroyed. List two ways thermal energy can be increased in a system. Adding thermal energy. Performing work on the system. Define the Second Law of Thermodynamics.

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Thermal energy flows from hot to cold. Define entropy.

Activity 1.3.3

Thermodynamics

Answer Key

Factual

thermodynamics

depends on the

essential

supposition that

every

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conceivable
setup of a given
framework, which
fulfill the
given limit
conditions, for
example,
temperature,
volume and
number of
particles, are
similarly prone
to happen. The
general

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framework will
in this manner
be in the
factually most
plausible setup.

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Properties 1. If
an object has a
weight of 10 lbf
on the moon,
what would the
same object
weigh on
Jupiter? Jupiter
22Moon c...

*Thermodynamic
Properties*
The First Law of
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Thermodynamics
Work and heat
are two ways of
transferring
energy between a
system and the
environment,
causing the
system's energy
to change. If
the system as a
whole is at
rest, so that
the bulk

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mechanical
energy due to
translational or
rotational
motion is zero,
then the

Chapter 17.

*Work, Heat, and
the First Law of
Thermodynamics*

Only the May
diet are
published,

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December diet questions are similar to questions A1-3 and section B in those papers. From 2019, the exam format is changed so that you need only answer ALL Section A (1-3 Thermodynamics, 4-6 statistical

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mechanics), ONE
question from
section B
(Thermodynamics)
and ONE from
section C
(Statistical
mechanics).

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The second law states that a process is spontaneous if the system and the surroundings have an increase in entropy.

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collection of
exercise
problems that
have been part

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