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electrons between atoms
is called a covalent
bond, and the two
electrons that join atoms
in a covalent bond are

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called a bonding pair of electrons. A discrete group of atoms connected by covalent bonds is called a

Covalent Bonding

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Metallic Bonds and

Metallic Properties

When a metal is

subjected to pressure,

the metal cations easily

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slide past one another. If an ionic crystal is struck with a hammer, the blow tends to push the positive ions close together.

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bonds are formed.

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Types of Bonds Ionic
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Bonds Bonds that are
formed by transfer of
electrons from one
element to the other.

Each element (now an
ion) will have a
complete octet after the
transfer of electrons.

The Ionic Bond □ The
electrical force of
attraction between
oppositely charged ions.

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There are three major types of chemical bonding. They are the ionic bonding, covalent bonding and metallic bonding. The key difference between ionic bonding and metallic bonding is that the ionic bonding takes place between positive and negative ions whereas

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the metallic bonding takes place between positive ions and electrons. Reference: 1. Libretexts. □ Ionic and Covalent Bonds. □ Chemistry LibreTexts, National Science Foundation, 28 Feb. 2018. Available here 2.

Difference Between
Ionic Bonding and
Metallic Bonding ...

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Answers Pearson

chemical bonds. There
are three types of
chemical bonds: ionic
bonding, covalent
bonding and metallic
bonding. This quiz will
focus on ionic bonding.
Use your knowledge of
ions and ionic bonding
to answer the following
questions. Group:

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Bonding : Chemical

Bonding I: Ionic

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BONDING AND INTERACTIONS. 7.1 Ions.

Essential Understanding
Ions form when atoms
gain or lose valence
electrons, becoming
electrically charged.

Lesson Summary.

Valence

Electrons Valence

electrons are the

electrons in the

outermost occupied

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energy level and are involved in ion formation. For a representative element, the group number equals the ...

BONDING AND INTERACTIONS

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Answer Key

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Group . 1A : 2A . 3A :
4A . 5A : 6A . 7A : 8A .
1 : 2 . 3 : 4 □ Notice that
all the electrons within a
given group (with the
exception of helium)

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have the same ...

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SECTION 7.2 IONIC
BONDS AND IONIC

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COMPOUNDS (pages 194–199) This section lists the characteristics of an ionic bond. It also describes the use of these characteristics to explain the electrical conductivity of ionic compounds when melted and when in aqueous solutions.

Formation of Ionic Compounds (pages 194–195) 1. What is an

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ionic bond? 2.

Bonding

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_____ 12. The coordination number of an ion is the number of ions of positive charge that surround the ion in a crystal. _____ 13. The coordination number of the ion Na^+ in NaCl is 6. _____ 14. In forming an ionic compound, an atom of an element gains electrons. _____ 15.

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IONS 7

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A covalent bond in
which three pairs of
electrons are shared by

two atoms. polyatomic

Ion. A group of

covalently bonded

atoms ... Chapter 3 □

Atoms and Ionic Bonds.

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Chapter 4 □ Atoms and
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metallic crystal
structures. Metallic
Bonds and Metallic

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Properties (pages

201–202) 1. Is the

following sentence true
or false? Metals are

made up of cations, not
neutral atoms. _____ 2.

What are metallic

bonds? 3. Name three

properties of metals that
can be explained by

metallic bonding. a. b. c.

4. What happens to an
ionic crystal when a ...

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SECTION 7.1 IONS

(pages 187–193)

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Formation of Cations

Group 2A Cations

Magnesium (atomic
number 12) belongs ...

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