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mole fraction = moles of solute / (moles of solute + moles of solvent) mole fraction equals the number of moles of solute or solvent in a solution divided by the number of moles of solute and solvent. mole fraction $x_A = n_A / (n_A + n_B)$ n is the number of moles of each substance. Concentrations can be measured.

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14. Possible answers include: flesh color (student. specific), solid filled with liquids, 7 cm, moves. through muscle and ligament connections. 15. Possible answers include: steel, easily magnetized, 3 1/4 cm × 3/4 cm, silver, solid. Section 2 (page 41) 1. physical. 2. chemical. 3. chemical. 4. physical. 5. chemical. 6. chemical. 7. chemical. 8. physical. 9. chemical

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1. Explain Use the properties of seawater to describe the characteristics of mixtures. Answers will vary but might include that seawater is a heterogeneous mixture with dirt and mud particles, and it is a homogeneous mixture with dissolved substances. 2.

Mixtures and Solutions

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