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Chapter 10 Chemical Quantities Study

Chemistry Chapter 10 Chemical Quantities. STUDY. PLAY. Mole. (mol) of a substance is 6.02×10^{23} representative particles of that substance and is the SI unit for measuring the amount of a substance. Avogadro's Number. the number of representative particles in a mole, 6.02×10^{23} . Representative Particle.

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Chapter 10 Chemical Quantities. mole. Avogadro's number. representative particle. molar mass. the SI unit for measuring the amount of a substance. number of representative particles in a mole, 6.02×10^{23} . refers to the species present in a substance: usually atoms, m.... the mass of one mole of a pure substance.

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Name _____ Date _____ Class _____ Chapter 10 – Chemical Quantities Avogadro's Number

One important property of a mole is that it means a definite number of particles just like a dozen means a number of particles. While a dozen is only 12 particles a mole is a larger number – 6.02×10^{23} particles.

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Chapter 10 Chemical Quantities91 SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance. Measuring Matter (pages 287–289) 1.

SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296)

Chemical Quantities THE MOLE AND QUANTIFYING MATTER 10.1 The Mole: A Measurement of Matter EssentialUnderstanding The mole represents a large number of very

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small particles.

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CHAPTER 10: Chemical Quantities BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance • The mole was founded by a scientist named Avagadro, and he

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