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1. Write the chemical and

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oxidation number of the positive ion. If the positive ion is monoatomic (only one atom), you can find its oxidation number from the periodic table. If the positive ion is polyatomic, use table 19.2 to find the oxidation number of the polyatomic ion.

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3.2 Composition of Compounds

A chemical formula is used to express the structure of a molecule. The formula tells which elements and how many of each element are present in a compound.

Formulas are written using the elemental symbol of each atom and a subscript to denote the number of elements.

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Ionic bonds, covalent bonds and metallic bonds are examples of chemical bonds. The structure and bonding in a substance are modelled in different ways, including dot and cross diagrams.

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element valence electrons

number needed element

valence electrons number

needed H11 Ne80 He20 Na17

Li17 Mg26 Be26 Al35 B35 Si44

C44 P53 N53 S62 O62 Cl71 F71

Ar80 Figure 19.3: Helium

atoms have only two

electrons, both of which are

in the outermost level.

Helium is an exception to

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Changes in Matter Chapter 19 Molecules and

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an atom that has an
electrical charge. ionic
compound A compound made out
of ions. Oxidation number
Shows how many electrons are
lost, Gained or

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